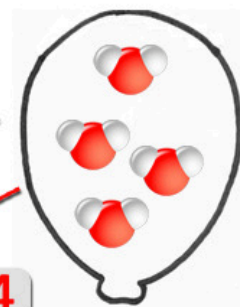




Avogadro (1776-1856)

MOLE (mol)

Volume gazeux (L)



$\times 22,4$

$: 22,4$

$: 6,02 \cdot 10^{23}$

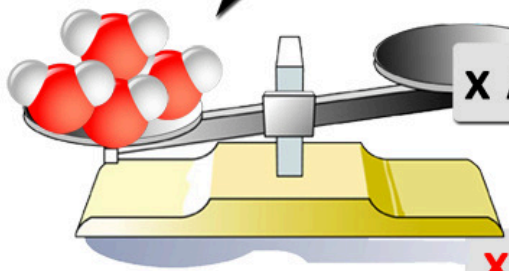
$\times 6,02 \cdot 10^{23}$

$\times Ar$ (ou Mr)

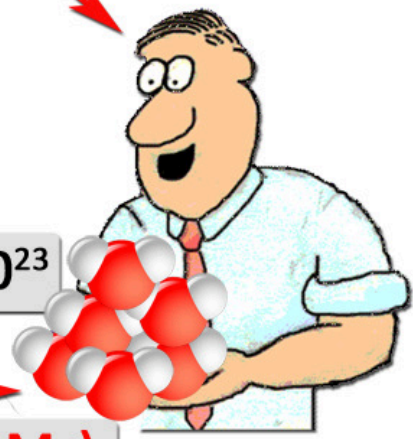
$: Ar$ (ou Mr)

$\times Ar$ (ou Mr) / $6,02 \cdot 10^{23}$

$\times 6,02 \cdot 10^{23} / Ar$ (ou Mr)



MASSE (g)



NOMBRE DE PARTICULES (n)